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A Direct Determination of the Molecular Structure of Prostaglandin F₂₋₁

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The tri-*p*-bromobenzoate of the methyl ester of prostaglandin F₂₋₁ (C₄₂H₄₇O₈Br₃) is orthorhombic with $a = 26.14$, $b = 33.93$ and $c = 4.76$ Å. The space group is $P2_12_12_1$ with $Z = 4$. The structure was solved by heavy-atom methods and refined by three-dimensional anisotropic least-squares treatment to give the molecular structure and configuration of prostaglandin F₂₋₁, which belongs to a new class of hormones having smooth-muscle-stimulating and blood-pressure-depressing activity.

Introduction

A smooth-muscle-stimulating and blood-pressure-depressing activity was discovered in human semen by Goldblatt (1933) and by von Euler (1934, 1935) who also found a similar activity in sheep sperm and named the factor 'prostaglandin'.

In 1960 Bergström & Sjövall (1960) isolated two active crystalline compounds PGE₁ (C₂₀H₃₄O₅) and PGF₁₋₁ (C₂₀H₃₆O₅) from sheep prostate glands. Reduction of PGE₁ with borohydride (Bergström *et al.*, 1962) yielded PGF₁₋₁ together with the isomeric compound PGF₂₋₁.

The tri-*p*-bromo- and tri-*p*-iodobenzoates of the methyl ester of PGF₂₋₁ were prepared by Prof. S. Bergström who kindly provided them for an X-ray analysis in order to determine *ab initio* the molecular structures.

Experimental

The two benzoates crystallize as beautiful needles from ethanol. All crystals studied gave excellent X-ray photographs, which indicated that the two derivatives are isomorphous. The unit cell is orthorhombic with space group $P2_12_12_1$. The cell dimensions are:

	<i>a</i>	<i>b</i>	<i>c</i>
Br-derivative	26.14 Å	33.93 Å	4.76 Å
I-derivative	26.49	34.39	4.76

Full three-dimensional data (Cu K α radiation) were collected for the bromo compound. Multiple-film Weissenberg photographs were recorded for the 0-2 layers about *c* (needle direction) and the 0-5 layers about *a*. Only *hk0* intensities were estimated for the iodo derivative. The intensities were corrected for the Lorentz and polarization factors but not for absorption.

Structure determination

The initial work was done on the (001) projection. From a sharpened Patterson projection of the I-derivative, the relative positions of the heavy atoms of the four molecules per cell were obtained by simple vector addition and subtraction. The origin was then determined from the known symmetry of the projection. As expected the Patterson synthesis was completely dominated by I-I vectors (Fig. 1).

An electron-density projection showed most of the atoms of the molecule (Fig. 2) but also contained some spurious detail. This, however, was still present,

when more atoms, *e.g.* those of the benzoate groups were included in the analysis and the general appearance of the maps did not improve much from that of Fig. 2. Similar effects were encountered with the Br-derivative $hk0$ data. They were caused by the large anisotropic vibrations of the heavy atoms. As soon as these were accounted for a projection could be found which was chemically sensible and which did refine. Anisotropic least-squares treatment of the Br-derivative projection data reduced R to 20%.

At that stage the three-dimensional data (1776

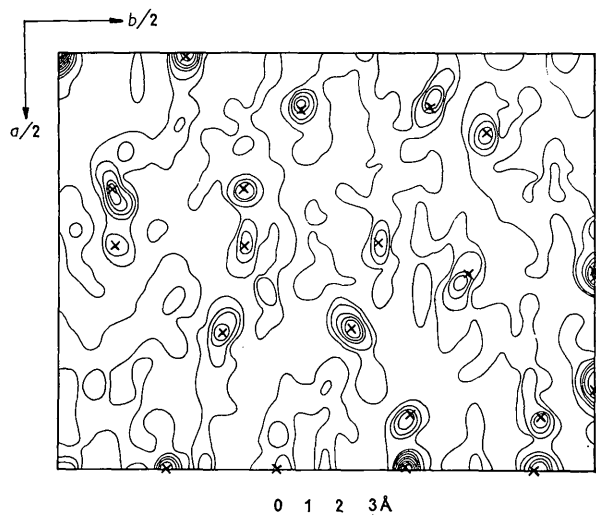


Fig. 1. Patterson (001) projection for the I-derivative. Crosses mark heavy atom vectors calculated from the final bromine positions.

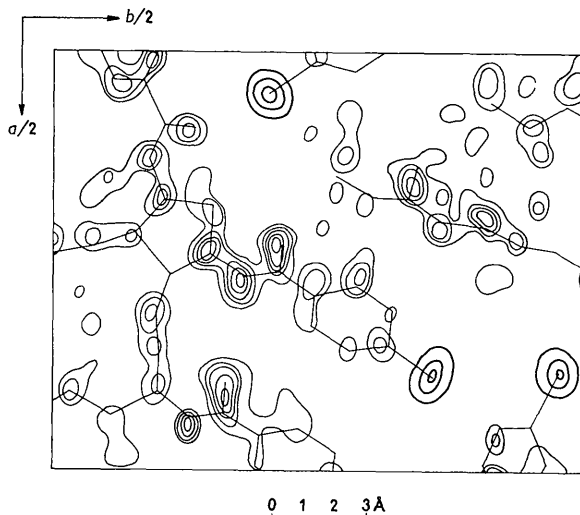


Fig. 2. Iodine-phased electron density projection along the c -axis with the correct structure outlined. Contours at arbitrary but equal levels except for the heavy atom, which has different contour intervals.

independent reflexions) were available. A Patterson synthesis sharpened to point atoms at rest (Abrahams & Maslen, 1963) was calculated. The resolution in the c -direction was still low due to the lack of high orders of l . This, and the presence of mirror planes, caused some trouble in finding the z -coordinates of the heavy atoms (see Fig. 3). Coordinates chosen to give the best compromise solution to the Patterson later proved to be surprisingly accurate (within 0.17 Å).

Table 1. Atomic coordinates

	x/a	y/b	z/c		x/a	y/b	z/c
Br ₁	0.63174	0.47212	0.01758	C ₁₇	0.59564	0.30890	-0.10939
Br ₂	0.63476	0.35299	-0.21371	C ₁₈	0.60688	0.27266	-0.24344
Br ₃	0.29836	0.19886	0.16054	C ₁₉	0.57848	0.23938	-0.15468
O ₁	0.34242	0.12472	-0.51624	C ₂₀	0.47724	0.03968	-0.28187
O ₂	0.37917	0.08772	-0.17795	C ₂₁	0.49046	0.00396	-0.08600
O ₃	0.51648	0.17763	-0.03904	C ₂₂	0.49111	0.47002	-0.23930
O ₄	0.47548	0.21145	0.27676	C ₂₃	0.48345	0.43217	-0.40866
O ₅	0.39422	0.33494	-0.06161	C ₂₄	0.45817	0.39924	-0.22800
O ₆	0.42681	0.29515	-0.37597	C ₂₅	0.45393	0.36086	-0.41734
O ₇	0.64002	0.16036	0.40039	C ₂₆	0.42221	0.33257	-0.23383
O ₈	0.68815	0.12388	0.11804	C ₂₇	0.39942	0.26075	-0.25072
C ₁	0.19407	0.05198	-0.11009	C ₂₈	0.55798	0.09409	0.09676
C ₂	0.19771	0.08252	-0.28724	C ₂₉	0.60572	0.09796	-0.01240
C ₃	0.24290	0.09588	-0.37575	C ₃₀	0.65203	0.09038	0.15999
C ₄	0.28700	0.08330	-0.24648	C ₃₁	0.67528	0.15801	0.24266
C ₅	0.28642	0.05488	-0.04880	C ₃₂	0.70630	0.19000	0.13500
C ₆	0.23859	0.03885	0.02603	C ₃₃	0.69792	0.22707	0.24057
C ₇	0.33866	0.10228	-0.32869	C ₃₄	0.72300	0.26250	0.14650
C ₈	0.42725	0.10450	-0.23528	C ₃₅	0.76164	0.25635	-0.03000
C ₉	0.43282	0.14464	-0.10760	C ₃₆	0.77360	0.21976	-0.14132
C ₁₀	0.48675	0.14279	0.04589	C ₃₇	0.74627	0.18892	-0.03900
C ₁₁	0.51282	0.10674	-0.10328	C ₃₈	0.68358	0.05368	0.07943
C ₁₂	0.46710	0.07703	-0.12347	C ₃₉	0.65506	0.01565	0.14440
C ₁₃	0.50569	0.20957	0.10708	C ₄₀	0.31319	0.47755	-0.56601
C ₁₄	0.54020	0.24402	0.02812	C ₄₁	0.33840	0.44075	-0.64926
C ₁₅	0.52861	0.28079	0.12427	C ₄₂	0.30974	0.40358	-0.57723
C ₁₆	0.55678	0.31405	0.10016				

Table 2. *Vibrational parameters*Allowance was made for anisotropic vibrations with $2-(h^2b_{11}+k^2b_{22}+l^2b_{33}+klb_{23}+hkb_{13}+hklb_{12})$

	b_{11}	b_{22}	b_{33}	b_{23}	b_{13}	b_{12}
Br ₁	0.00293	0.00392	0.29367	-0.00325	-0.01480	0.00287
Br ₂	0.00878	0.00236	0.28466	0.00437	0.00991	-0.00608
Br ₃	0.00431	0.00167	0.24959	0.00221	0.00609	0.00184
O ₁	0.00262	0.00154	0.16485	0.00902	0.00414	0.00097
O ₂	0.00241	0.00158	0.04023	0.00450	-0.00858	-0.00018
O ₃	0.00208	0.00095	0.10525	-0.00587	0.00649	-0.00086
O ₄	0.00521	0.00114	0.19606	-0.00059	0.00333	-0.00070
O ₅	0.00409	0.00170	0.15280	0.00683	0.00976	-0.00251
O ₆	0.00555	0.00149	0.17847	0.00366	0.00426	0.00074
O ₇	0.00406	0.00239	0.23268	-0.00167	0.01554	-0.00369
O ₈	0.00220	0.00206	0.03925	0.00169	0.00254	-0.00108
C ₁	0.00474	0.00172	0.18647	-0.00406	0.02437	-0.00206
C ₂	0.00161	0.00347	0.23723	-0.00858	-0.00886	-0.00059
C ₃	0.00329	0.00218	0.33806	0.00369	-0.00505	0.00178
C ₄	0.00345	0.00130	0.26335	-0.01230	-0.00462	-0.00110
C ₅	0.00294	0.00164	0.26255	-0.00572	-0.01158	-0.00188
C ₆	0.00263	0.00317	0.21271	0.01304	0.00274	-0.00199
C ₇	0.00197	0.00134	0.17372	0.01241	-0.00242	-0.00137
C ₈	0.00120	0.00103	0.08023	0.00371	0.01110	-0.00080
C ₉	0.00295	0.00093	0.14333	-0.00480	0.00034	-0.00115
C ₁₀	0.00185	0.00093	0.13316	-0.00518	0.00568	-0.00108
C ₁₁	0.00139	0.00074	0.20162	0.00628	0.00733	-0.00072
C ₁₂	0.00190	0.00081	0.17588	-0.01310	-0.00329	0.00010
C ₁₃	0.00338	0.00106	0.21594	0.00536	0.00169	0.00104
C ₁₄	0.00310	0.00135	0.07704	-0.00545	0.00767	-0.00304
C ₁₅	0.00427	0.00115	0.30794	0.00247	0.00390	-0.00095
C ₁₆	0.00502	0.00186	0.06394	-0.00083	-0.00206	-0.00248
C ₁₇	0.00639	0.00152	0.19506	0.01874	0.01861	-0.00129
C ₁₈	0.00512	0.00105	0.23684	-0.00604	0.01575	-0.00250
C ₁₉	0.00514	0.00081	0.22429	-0.00166	0.04025	-0.00079
C ₂₀	0.00287	0.00084	0.13278	0.00140	-0.01379	-0.00038
C ₂₁	0.00334	0.00090	0.24077	-0.00104	-0.02731	0.00151
C ₂₂	0.00400	0.00062	0.11502	-0.00293	-0.00679	0.00049
C ₂₃	0.00304	0.00106	0.12249	-0.00260	-0.00228	-0.00114
C ₂₄	0.00356	0.00119	0.12110	-0.01377	-0.01040	-0.00047
C ₂₅	0.00341	0.00038	0.26662	0.00114	0.00778	-0.00088
C ₂₆	0.00276	0.00102	0.08223	-0.00112	0.01062	0.00041
C ₂₇	0.00396	0.00096	0.20239	0.00053	0.01334	-0.00261
C ₂₈	0.00246	0.00022	0.15666	0.00024	-0.00348	-0.00004
C ₂₉	0.00212	0.00092	0.19003	0.00467	0.01987	-0.00008
C ₃₀	0.00240	0.00057	0.15615	-0.00493	-0.00688	-0.00004
C ₃₁	0.00252	0.00094	0.20924	0.00691	-0.02245	-0.00128
C ₃₂	0.00331	0.00140	0.14783	0.00240	0.01644	-0.00053
C ₃₃	0.00487	0.00236	0.28037	-0.00400	0.02831	-0.00182
C ₃₄	0.00581	0.00106	0.26741	0.01266	0.03284	-0.00085
C ₃₅	0.00363	0.00038	0.25971	-0.00464	0.00824	-0.00082
C ₃₆	0.00336	0.00146	0.28297	-0.00609	0.01990	-0.00405
C ₃₇	0.00273	0.00185	0.32248	-0.00107	0.02743	0.00014
C ₃₈	0.00345	0.00120	0.10454	0.00545	-0.01850	-0.00076
C ₃₉	0.00400	0.00033	0.19997	-0.00638	0.02130	0.00032
C ₄₀	0.00261	0.00021	0.19404	-0.00685	-0.02013	0.00009
C ₄₁	0.00350	0.00129	0.21235	-0.00036	0.00382	0.00145
C ₄₂	0.00468	0.00113	0.20270	0.00345	-0.02783	-0.00103

The three-dimensional electron-density series calculated on the bromine phases showed the complete molecule as found from the (001) projection. Two rounds of Fourier refinement, during which least-squares methods were applied to the heavy atoms, reduced R to 23% for all observed reflexions and confirmed the assignment of atomic numbers to the peaks in the projection analysis.

All atoms were finally included in the anisotropic least-squares refinement. Three cycles reduced R to 10.9%. It was not considered worth while to perform further 7-hour least-squares cycles as the molecular

structure is fully established with the accuracy in positions reached. However, the shifts in the least accurately placed atoms (the benzene-ring carbons) approached the standard deviations.

For carbon and oxygen the scattering curves of Berghuis *et al.* (1955) were used and for bromine that of Thomas & Umeda (1957) corrected for the f' component of the anomalous dispersion (Dauben & Templeton, 1955). The calculations were performed on a Ferranti-Mercury computer with programs described by Mills & Rollett (1961). The weighting scheme adopted in the least-squares refinement was

Table 3. *Interatomic distances and bond angles*

Bond distances			
Bond	Bond distance	Bond	
Br ₁ -C ₁	1.88 Å	C ₃₃ -C ₃₄	1.44 Å
Br ₂ -C ₁₇	1.88	C ₃₄ -C ₃₅	1.33
Br ₃ -C ₃₅	1.90	C ₃₅ -C ₃₆	1.39
C ₁ -C ₂	1.34	C ₃₆ -C ₃₇	1.36
C ₂ -C ₃	1.33	C ₃₇ -C ₃₂	1.33
C ₃ -C ₄	1.37	C ₈ -C ₉	1.50
C ₄ -C ₅	1.35	C ₉ -C ₁₀	1.59
C ₅ -C ₆	1.41	C ₁₀ -C ₁₁	1.57
C ₇ -C ₆	1.40	C ₁₁ -C ₁₂	1.57
C ₄ -C ₇	1.55	C ₁₂ -C ₂₀	1.50
C ₇ -O ₁	1.18	C ₂₀ -C ₂₁	1.58
C ₇ -O ₂	1.37	C ₂₁ -C ₂₂	1.52
O ₂ -C ₈	1.41	C ₂₂ -C ₂₃	1.53
C ₁₀ -O ₃	1.47	C ₂₃ -C ₂₄	1.56
O ₃ -C ₁₃	1.32	C ₂₄ -C ₂₅	1.59
C ₁₃ -O ₄	1.13	C ₂₅ -C ₂₆	1.54
C ₁₃ -C ₁₄	1.52	C ₂₆ -O ₅	1.10
C ₁₄ -C ₁₅	1.36	C ₂₆ -O ₆	1.44
C ₁₅ -C ₁₆	1.35	O ₆ -C ₂₇	1.49
C ₁₆ -C ₁₇	1.43	C ₁₁ -C ₂₈	1.58
C ₁₇ -C ₁₈	1.42	C ₂₈ -C ₂₉	1.36
C ₁₈ -C ₁₉	1.42	C ₂₉ -C ₃₀	1.48
C ₁₄ -C ₁₉	1.34	C ₃₀ -C ₃₈	1.54
C ₃₀ -O ₈	1.49	C ₃₈ -C ₃₉	1.52
O ₈ -C ₃₁	1.34	C ₃₉ -C ₄₀	1.58
O ₇ -C ₃₁	1.19	C ₄₀ -C ₄₁	1.47
C ₃₁ -C ₃₂	1.45	C ₄₁ -C ₄₂	1.51
C ₃₂ -C ₃₃	1.37		

Bond angles			
Bond angle	Bond angle	Bond angle	
Br ₁ -C ₁ -C ₂	123°	C ₁₀ -C ₁₁ -C ₂₈	105°
Br ₁ -C ₁ -C ₆	118	C ₁₂ -C ₁₁ -C ₂₈	116
C ₆ -C ₁ -C ₂	118	C ₁₁ -C ₁₂ -C ₈	99
C ₁ -C ₂ -C ₃	122	C ₁₁ -C ₁₂ -C ₂₀	116
C ₂ -C ₃ -C ₄	120	C ₈ -C ₁₂ -C ₂₀	118
C ₃ -C ₄ -C ₅	122	C ₁₂ -C ₂₀ -C ₂₁	112
C ₂ -C ₄ -C ₃	119	C ₂₀ -C ₂₁ -C ₂₂	108
C ₂ -C ₄ -C ₅	119	C ₂₁ -C ₂₂ -C ₂₃	112
C ₄ -C ₅ -C ₆	118	C ₂₂ -C ₂₃ -C ₂₄	112
C ₅ -C ₆ -C ₁	120	C ₂₃ -C ₂₄ -C ₂₅	108
O ₁ -C ₇ -C ₄	122	C ₂₄ -C ₂₅ -C ₂₆	103
O ₂ -C ₇ -C ₄	113	C ₂₅ -C ₂₆ -O ₅	137
O ₁ -C ₇ -O ₂	124	C ₂₅ -C ₂₆ -O ₆	104
C ₇ -O ₂ -C ₈	116	O ₅ -C ₂₆ -O ₆	118
C ₁₀ -O ₃ -C ₁₃	114	C ₂₆ -O ₆ -C ₂₇	117
O ₃ -C ₁₃ -O ₄	125	C ₁₁ -C ₂₈ -C ₂₉	116
O ₃ -C ₁₃ -C ₁₄	112	C ₂₈ -C ₂₉ -C ₃₀	121
O ₄ -C ₁₃ -C ₁₄	123	C ₂₉ -C ₃₀ -C ₃₈	116
C ₁₃ -C ₁₄ -C ₁₅	119	C ₂₉ -C ₃₀ -O ₈	108
C ₁₃ -C ₁₄ -C ₁₉	120	O ₈ -C ₃₀ -C ₃₈	104
C ₁₉ -C ₁₄ -C ₁₅	121	C ₃₀ -C ₃₈ -C ₃₉	112
C ₁₄ -C ₁₅ -C ₁₆	128	C ₃₈ -C ₃₉ -C ₄₀	113
C ₁₅ -C ₁₆ -C ₁₇	110	C ₃₉ -C ₄₀ -C ₄₁	113
C ₁₆ -C ₁₇ -C ₁₈	124	C ₄₀ -C ₄₁ -C ₄₂	115
C ₁₆ -C ₁₇ -Br ₂	118		
C ₁₈ -C ₁₇ -Br ₂	117		
C ₁₈ -C ₁₉ -C ₁₄	120		
C ₃₀ -O ₈ -C ₃₁	116		
O ₈ -C ₃₁ -O ₇	122		
O ₈ -C ₃₁ -C ₃₂	110		
O ₇ -C ₃₁ -C ₃₂	127		
C ₃₁ -C ₃₂ -C ₃₃	118		
C ₃₁ -C ₃₂ -C ₃₇	130		
C ₃₇ -C ₃₂ -C ₃₃	112		
C ₃₂ -C ₃₃ -C ₃₄	125		
C ₃₃ -C ₃₄ -C ₃₅	114		
C ₃₄ -C ₃₅ -C ₃₆	124		
C ₃₄ -C ₃₅ -Br ₃	118		
C ₃₆ -C ₃₅ -Br ₃	119		
C ₃₅ -C ₃₆ -C ₃₇	116		
C ₃₆ -C ₃₇ -C ₃₂	128		
C ₁₂ -C ₈ -C ₉	111		
O ₂ -C ₈ -C ₉	112		
O ₂ -C ₈ -C ₁₂	108		
C ₈ -C ₉ -C ₁₀	104		
C ₉ -C ₁₀ -C ₁₁	102		
C ₉ -C ₁₀ -O ₃	108		
C ₁₁ -C ₁₀ -O ₃	106		
C ₁₀ -C ₁₁ -C ₁₉	101		

$$w = \frac{1}{1 + [(|F_o| - 8|F_{\min.}|)/5|F_{\min.}|]^2}$$

Discussion

The coordinates of the atoms in the molecule other than hydrogen are given in Table 1, the vibrational parameters in Table 2. The final structure factors are listed in Table 6. A composite drawing of the three-

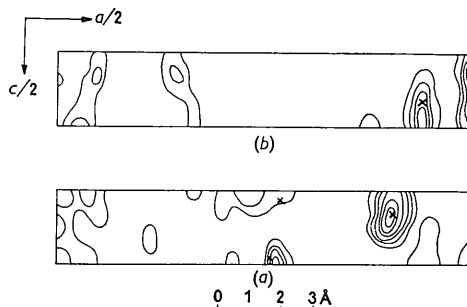


Fig. 3. Sections of the three-dimensional point atoms at rest. Patterson synthesis for the bromo derivative. (a) Harker section at $z=1/2$, (b) section at $z=26/80$. Crosses mark Br-Br vectors calculated from the final positions.

dimensional electron-density series based on the final phases is shown in Fig. 4.

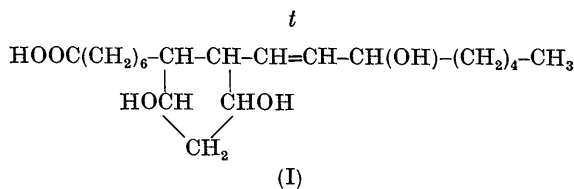
Table 4. *Average values for different classes of bond distances and angles*

	Bond		Angle
Benzene rings	C-C	1.38 ± 0.05 Å	$120 \pm 10^\circ$
Unstrained carbon single bonds	C-C	1.54 ± 0.07	111 ± 8
Carboxyl carbon to oxygen	C=O	1.15 ± 0.05	
	C-O	1.37 ± 0.07	
Other carbon-oxygen	C-O	1.46 ± 0.05	
	Br-C	1.89 ± 0.01	

Table 5. *Some short contacts between one molecule and its equivalent repeated along the c-axis*

O ₁ -O ₂	3.52 Å	C ₁₅ -C ₁₈	3.64 Å
O ₃ -O ₄	3.62	C ₁₅ -C ₁₉	3.93
O ₇ -O ₈	3.84	C ₁₆ -C ₁₈	3.67
O ₅ -O ₆	3.63	C ₃₃ -C ₃₆	3.55
C ₂ -C ₆	3.74	C ₃₃ -C ₃₇	3.88
C ₃ -C ₅	3.67	C ₃₄ -C ₃₆	3.92
C ₈ -C ₆	3.44	C ₈ -C ₁₀	3.98

The molecular structure of the tri-*p*-bromobenzoate derivative of the methyl ester of PGF₂₋₁(C₄₂H₄₇O₅Br₃) is all in accordance with the formula (I) for the free acid of PGF₂₋₁ independently arrived at by Bergström *et al.* (1962a) using chemical and physicochemical methods other than X-ray diffraction.



PGF₂₋₁ is thus a trihydroxy acid with two hydroxyl groups attached directly to a five membered hydrocarbon ring and the third along a carbon side chain

Table 6 (cont.)

Table with multiple columns of numerical data, likely representing bond lengths, angles, or vibrational frequencies for prostaglandin F2-1. The table is organized into several vertical columns and rows, with some entries containing '+' or '-' signs, possibly indicating phase or direction.

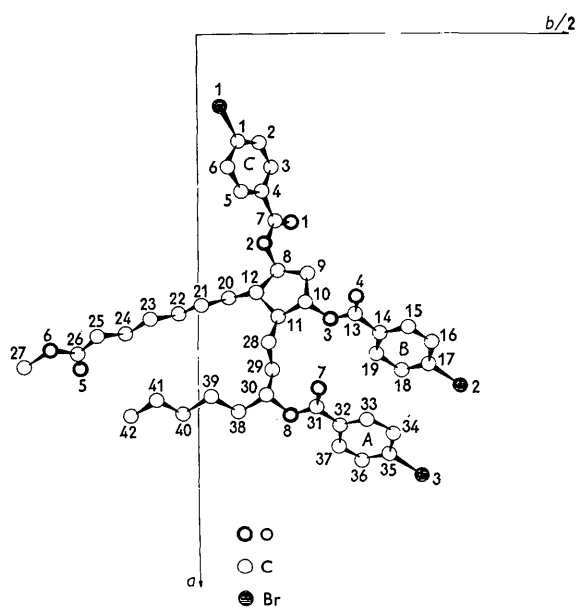


Fig. 5. Spatial diagram of the molecule (correct absolute configuration) with the numbering of atoms indicated.

slightly better if one compares distances and angles with the corresponding generally accepted values but

is, of course, still low in a heavily vibrating structure of this complexity. The average values for different classes of distances and angles are given in Table 4.

The molecular packing is illustrated in Fig. 6. The side chains and the benzoate groups project out from the five-membered ring to give a molecule extended in the (001) plane. The two side chains and the bromobenzoate groups *A* and *B* (Fig. 5) are in close contact whereas other molecules can approach the five-membered ring on both sides of group *C*. Similar parts of the molecules are in contact along the short *c*-axis, but this is not possible in other directions and carbon chains have to pack with benzene rings. However, the carbon chains group together as far as possible and alternate with regions of bromobenzoate groups. Some of the most important contacts are shown in Fig. 6 whereas contacts between one molecule and the molecule directly above or below in the *c*-direction are given in Table 5.

The benzene ring, the bromine atom and the carboxyl carbon are in all bromobenzoate groups planar within 0.08 Å. There is an indication that the carboxyl oxygen atoms may be slightly twisted out of the plane (up to 0.15 Å). Benzoate groups *A* and *B* are approximately parallel. Their planes make an angle of 42° with the (001) plane and their twofold axes are tilted 15° to the same plane. The third benzoate

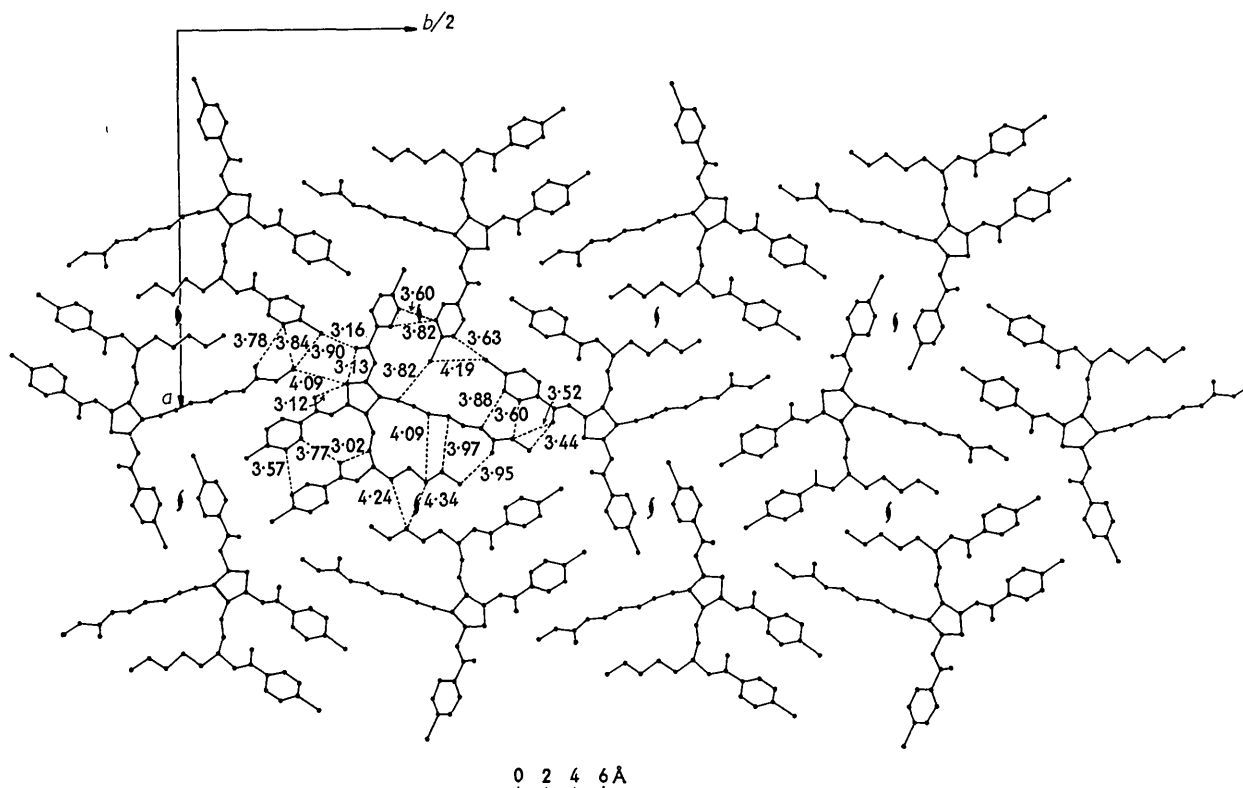


Fig. 6. Molecular arrangement of the *p*-bromobenzoate derivative of the methyl ester of PGF₂₋₁ as seen along the *c*-axis with some of the more important contacts shown.

group shows very similar angles but has its diad axis roughly perpendicular to those of the other two groups when the three groups are in packing contact. Most benzene rings in this structure pack with their planes parallel but *C*-rings also make contact with roughly perpendicular planes at screw axes.

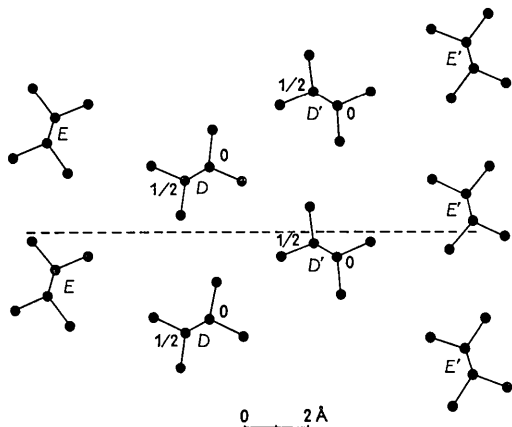


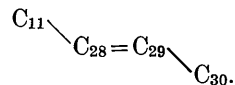
Fig. 7. Schematic representation of the hydrocarbon chain packing as seen along the chain axes. 0 and 1/2 are fractions of the repeat distance (2.55 Å) along a chain. Dotted line indicates the plane $z=0$.

The hydrocarbon chains are surprisingly regular considering their small length and their packing contacts with benzene rings. Atoms C₂₁ to C₂₅ are planar within 0.05 Å whereas C₂₀ is 0.16 Å off the plane. C₃₉ to C₄₂ lie within 0.002 Å from a plane with C₃₈ at a distance of 0.12 Å. The axes of the two chains are roughly parallel to the (001) plane but are not mutually parallel as the position of the side chains in the molecule prevents them coming into close-packing contact until near their ends. The carbon-chain arrangement is illustrated schematically in Fig. 7. The plane of chain *D* is inclined 28° to the (001) plane and that of chain *E* 76° in the same sense. It is not to be expected that the chain packing should be exactly similar to those found in long-chain structures (cf. Abrahamsson *et al.*, 1963) but the general arrangement of *D*-chains resembles very much that of the common orthorhombic type O₁.

The atoms of the ester group (C₂₅, C₂₆, O₅, O₆, C₂₇) lie within 0.06 Å or less from a plane, which is twisted 23° from that of the adjoining hydrocarbon chain. A similar angle of twist (19°) is found in methyl stearate (Aleby & von Sydow, 1960).

Four of the atoms (C₈, C₉, C₁₀, C₁₂) of the five-membered ring lie in a plane within 0.07 Å whereas the fifth is at an appreciable distance (0.71 Å) as is usual for such saturated rings. The bonds to the different attached groups are all *trans* relative to the

five membered ring (Fig. 5). The bonds C₃₀-O₈ and C₃₀-C₃₈ are rotated 45° and 54° respectively out of the plane of the double bonded group



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